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**COMPOSITION RANGE AND GLASS
FORMING ABILITY OF TERNARY
Ca-Mg-Cu BULK METALLIC
GLASSES (Preprint)**



O.N. Senkov, J.M. Scott, and D.B. Miracle

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Composition Range and Glass Forming Ability of Ternary Ca-Mg-Cu Bulk Metallic Glasses

O.N. Senkov ^{*1}, J.M. Scott*, and D.B. Miracle

Air Force Research Laboratory, Materials and Manufacturing Directorate, Wright-Patterson AFB, OH 45433, USA; *UES, Inc., 4401 Dayton-Xenia Road, Dayton, OH 45432, USA.

ABSTRACT

A number of ternary Ca-Mg-Cu amorphous alloys with compositions ranging from 40 to 70 at.% for Ca, 5 to 25 at. % for Mg and 10 to 35 at. % for Cu were produced by a copper mold casting method as wedge-shaped samples with the thickness varying from 0.5 mm to 10 mm. The maximum thickness at which an alloy remains fully amorphous, glass transition temperature, crystallization temperature, temperature interval of the super-cooled region, solidus and liquidus temperatures, as well as heats of crystallization and melting, are reported for these alloys. The effect of the alloy composition on glass forming ability is discussed. In the studied composition range, the Ca₅₀Mg_{22.5}Cu_{27.5} alloy has the best glass forming ability, being fully amorphous after casting to the thicknesses of at least 10 mm.

Keywords: Casting; Ca-Mg-Cu alloys; bulk metallic glass; differential scanning calorimetry; X-ray diffraction; glass forming ability.

¹ Corresponding author. Phone: 937-255-1320; Fax: 937-656-7292; E-mail: oleg.senkov@wpafb.af.mil

INTRODUCTION

Ca-based bulk metallic glasses are a relatively new class of amorphous alloys which have unique properties such as low density (~ 2.0 g/cc or less), low Young's modulus (~17-20 GPa), low glass transition temperature ($T_g \sim 80-150^\circ\text{C}$) and a wide temperature range of super-cooled liquid ($\Delta T_x = 30-80^\circ\text{C}$). With the exception of Mg-based alloys, this is the only known simple-metal based system in the group of bulk metallic glasses. The first Ca-based glassy alloys, $\text{Ca}_{57}\text{Mg}_{19}\text{Cu}_{24}$, $\text{Ca}_{60}\text{Mg}_{20}\text{Ag}_{20}$, and $\text{Ca}_{60}\text{Mg}_{20}\text{Ag}_{10}\text{Cu}_{10}$, were reported by Amiya and Inoue in 2002 [1,2]. Fully amorphous rods with a diameter of up to 4 mm in the case of the ternary alloys and up to 7 mm for the quaternary alloy were prepared by a copper mold casting method. To date, numerous Ca-based bulk-glass forming alloys have been reported in the literature that also include Ca-Mg-Zn [3-6], Ca-Mg-Al [3,7,8], Ca-Al-Cu [3,7,8], Ca-Al-Zn [8], Ca-Al-Ag [8], Ca-Mg-Zn-Cu [3,4], Ca-Y-Mg-Cu [3,7], Ca-Mg-Al-Zn [3,7], Ca-Mg-Al-Cu [7,8], Ca-Mg-Al-Ag-Cu [7], and others. The composition of all these alloys can be described by the formula [3]:



where A = 0.40 to 0.70; B = 0 to 0.25; C = 0 to 0.25; D = 0 to 0.35; E = 0 to 0.35; B+C+D ≥ 0.05; A+B+C+D+E=1, and Ln represents lanthanide group elements. The composition range described by Equation (1) was identified using the specific criteria for selection of bulk glass forming alloy compositions, which have recently been identified [3,7,9].

Although the $\text{Ca}_{57}\text{Mg}_{19}\text{Cu}_{24}$ was the first bulk metallic glass produced in the Ca-Mg-Cu system [1], no extensive work has been conducted yet to identify other bulk glassy alloys with better glass forming ability in this system. In the present work, a systematic study of glass forming ability in the ternary Ca-Mg-Cu alloy system in the composition range from 40 to 70 at.% for Ca, 5 to 25 at % for Mg and 10 to 35 at % for Cu was conducted and a number of new bulk glassy alloys were produced. Thermal properties of the new amorphous alloys are reported and the glass forming abilities are correlated with the alloy composition.

EXPERIMENTAL

$\text{Ca}_X\text{Mg}_Y\text{Cu}_Z$ alloys, where X = 40 to 70 at.%; Y = 5 to 30 at.%; and Z = 10 to 36 at.%, were prepared by induction melting of mixtures of pure elements (99.9%) in a water-cooled copper crucible in an argon atmosphere. Some of these alloys were also prepared by melting the elements in a closed-end quartz tube using a radiant furnace and argon atmosphere. Each of the

produced alloys was then induction re-melted in a quartz crucible with a 12.5 mm inner diameter and a 2-3 mm diameter hole in the bottom of the crucible and cast through this hole into a water-cooled copper mold to produce wedge-like samples. Wedge-shaped samples of two different dimensions were produced. One type of samples had 5 mm width, 30 mm height, and thickness varying from 0.5 mm at the base of the wedge to 3 mm at the top of the wedge. Another type of samples had 10 mm width, 50 mm height, and thickness varying from 2 mm at the base of the wedge to 10 mm at the top of the wedge. The time to fill the mold cavity was about 0.5 seconds for smaller samples and ~1.0-1.5 seconds for larger samples. The amorphous or crystalline structure of the cast alloys was identified by X-ray diffraction of powdered samples using Rigaku Rotaflex diffractometer, Cu K_α radiation, and by differential scanning calorimetry (DSC). Glass transition and crystallization temperatures, as well as heats of crystallization and melting, of the cast alloys were determined using a differential scanning calorimeter DSC Q1000 by TA Instruments Inc., (109 Lukens Drive, New Castle, DE 19720) at a heating rate of 40 K/min. The solidus and liquidus temperatures of the alloys were determined by DSC using heating and cooling cycles at 10°C/min. The weight of samples used for DCS was in the range of 6 to 15 mg.

RESULTS

Figure 1a illustrates typical DSC curves of a copper mold cast alloy Ca₅₃Mg₂₃Cu₂₄ of different thicknesses. An endothermic reaction corresponding to the glass transition into a super-cooled liquid condition and following exothermic reactions due to crystallization are characteristics of this alloy and indicate that the alloy contains an amorphous phase. At higher temperatures, an extensive endothermic reaction is seen due to melting. The heat of crystallization (ΔH_x) almost does not depend on the thickness of the cast specimen, τ , if the latter is below the critical value τ_{cr} ; however, ΔH_x rapidly decreases at larger thicknesses, Figure 1b. The magnitude of the heat of crystallization is known to have maximum when the sample is fully amorphous, and it decreases when the volume fraction of the amorphous phase decreases [10]. Therefore, the maximum value of the heat of crystallization, which remains constant at the thicknesses below τ_{cr} , is an indication that at these thicknesses the alloy is fully amorphous, and at the thicknesses above τ_{cr} the alloy becomes partially amorphous. X-ray diffraction (XRD) analysis supports this conclusion. Indeed, the XRD patterns from the samples extracted from regions of the wedge cast specimen with the thicknesses below τ_{cr} show a single wide halo with the 2-Theta range of 30°-

35° from the amorphous phase only, while the XRD patterns taken from the specimen regions having the thickness larger than τ_{cr} show also sharp crystalline peaks from crystalline phases (Figure 2).

Figure 3 shows (a) DSC and (b) XRD patterns of the cast alloys $Ca_{58}Mg_{18}Cu_{24}$ and $Ca_{50}Mg_{25}Cu_{25}$. The patterns are taken from the regions of the critical thickness τ_{cr} (which is 6 mm for $Ca_{58}Mg_{18}Cu_{24}$ and 9 mm for $Ca_{50}Mg_{25}Cu_{25}$) and slightly above τ_{cr} . One can see that the intensities of the exothermic crystallization reactions on the DSC curves decrease and sharp crystalline peaks are formed on the XRD patterns for both alloys, when the thickness exceeds τ_{cr} . The values of τ_{cr} for different Ca-Mg-Cu alloys studied in the present work are given in Table 1. Characteristic temperatures, such as the glass transition, T_g , onset of crystallization, T_x , the temperatures of the start, T_m , and completion, T_l , of melting, as well as heats of crystallization, ΔH_x , and melting (fusion), ΔH_m , are also tabulated in Table 1 for these alloys. New compositions have been identified with very good glass forming abilities. For example, the alloy $Ca_{50}Mg_{22.5}Cu_{27.5}$ is fully amorphous at a thickness of at least 10 mm, the alloy $Ca_{50}Mg_{25}Cu_{25}$ is fully amorphous up to a thickness of 9 mm, the alloy $Ca_{50}Mg_{20}Cu_{30}$ has the critical thickness of 8 mm, and the alloys $Ca_{45}Mg_{25}Cu_{30}$ and $Ca_{58}Mg_{18}Cu_{24}$ are fully amorphous up to the critical thickness of 6 mm.

Figure 4 illustrates (a) DSC and (b) XRD patterns from the cast alloys $Ca_{70-X}Mg_{20}Cu_{10+X}$ of the critical thickness, where $X = 0; 5; 10; 15$; and 20. The glass transition and crystallization temperatures are seen to increase with an increase in X , which is an indication that the amorphous phase, once it is formed, is kinetically more stable in the alloys with the larger amount of Cu (at a constant Mg content, see also Table 1). On the other hand, the glass forming ability (GFA), or critical thickness, does not increase monotonically with an increase in the amount of Cu, but it has intermediate maxima for $Ca_{60}Mg_{20}Cu_{20}$ ($\tau_{cr} = 4$ mm) and $Ca_{50}Mg_{20}Cu_{30}$ ($\tau_{cr} = 8$ mm), see Figure 4a and Table 1. This result may indicate that there is no direct correlation between the GFA and the glass stability. Analysis of the X-ray diffraction patterns shows that the position of the amorphous halo maximum shifts to larger 2Θ (from 29.5° to 32.3°) and the halo becomes wider (i.e. the full-width-at-half-maximum [FWHM] increases from $\Delta 2\Theta = 3.99^\circ$ to 9.06°) with an increase in Cu (from 10 to 30 at.%).

DISCUSSION

The results of the present work show that fully amorphous Ca-based ternary Ca-Mg-Cu alloys can be produced as plates with the thickness of 10 mm; which is much higher than a previously reported 4-mm maximum diameter rod of a fully amorphous ternary alloy $\text{Ca}_{57}\text{Mg}_{19}\text{Cu}_{24}$ [1]. The glass forming ability (GFA) of these ternary alloys is found to be very sensitive to the alloy composition. For example, the alloy $\text{Ca}_{50}\text{Mg}_{25}\text{Cu}_{25}$ has the maximum critical thickness of 9 mm; while the alloys $\text{Ca}_{50}\text{Mg}_{30}\text{Cu}_{20}$ and $\text{Ca}_{45}\text{Mg}_{30}\text{Cu}_{25}$, which compositions differ from the former alloy by only 5 at.%, have the critical thicknesses of 2 mm and 1 mm, respectively. Although the critical thickness depends on the casting conditions and it generally increases with increasing the cooling rate during solidification, in the case of the same casting procedure there is a direct relationship between the critical thickness and glass forming ability (GFA) [11], so that an alloy with larger τ_{cr} is a better glass former.

The composition range for the prospective Ca-Mg-Cu bulk glass forming alloys was selected in the present work using Equation (1), and this range is shown in the currently available liquidus projection of the Ca-Mg-Cu system (Figure 5) as a trapezoid-like area surrounded by a dashed-line border. The compositions of 27 amorphous alloys studied in the current work are shown in this figure as solid circles. The selected (trapezoid) composition area contains, according to [12], a ternary eutectic reaction and two peritectic reactions. The eutectic ($\sim\text{Ca}_{67}\text{Mg}_{19}\text{Cu}_{14}$) and one peritectic ($\sim\text{Ca}_{63}\text{Mg}_{19.5}\text{Cu}_{17.5}$) compositions are located in the Ca-rich region of this selected area, and the alloys in this region have rather moderate GFA, showing the critical thickness, τ_{cr} , of 4 mm or less (see Table 1). This agrees with previous results of Amiya and Inoue [1] who reported that the alloy $\text{Ca}_{67}\text{Mg}_{19}\text{Cu}_{14}$ can be copper mold cast into a fully amorphous rod with a maximum diameter of 2 mm only. The GFA has a tendency to increase when the concentration of Ca decreases to about 50 at.% and the amount of Mg is within 19-25 at.% range (see Table 1 and Figure 5); so that the alloys located near the second peritectic (the composition of this peritectic, according to [12], is $\sim\text{Ca}_{46}\text{Mg}_{25}\text{Cu}_{29}$) are fully amorphous at τ_{cr} of 6 to 10 mm. Such behavior of the Ca-Mg-Cu glasses is different from the behavior of Ca-Mg-Zn glassy alloys that have the best GFA (and τ_{cr} of ~ 10 mm) at a near eutectic composition [5], and it also contradicts the experimental observation that the best glass forming alloys have near-eutectic compositions [11].

Thorough analysis of the melting behavior of the Ca-Mg-Cu alloys (see Table 1) and results reported by Myles [12] has led us to a conclusion that this apparent contradiction is due to an incorrect assessment of the Ca-Mg-Cu liquidus projection made by Myles. The diagram shown in Figure 5 was constructed in [12] based on the metallographic analysis of 15 different alloys only. The phase identification was partially supported by X-ray diffraction and electron-probe analysis; however, no DSC was conducted to identify the melting reactions and liquidus temperatures [12]. For example, according to this diagram, the alloys $\text{Ca}_{70}\text{Mg}_{20}\text{Cu}_{10}$ and $\text{Ca}_{70}\text{Mg}_{10}\text{Cu}_{20}$ should experience the eutectic reaction at ~67 at.% Ca and 19% Mg and $T_m = 350 \pm 5^\circ\text{C}$; however, according to our results (see Table 1) these alloys have the highest melting temperatures among the all studied alloys, $T_m = 397^\circ\text{C}$ and $T_m = 386^\circ\text{C}$, respectively. The DSC analysis of 27 alloys conducted in the present work shows at least four characteristic temperatures for T_m (see Table 1). These are $T_m = 397^\circ\text{C}$ for the alloy $\text{Ca}_{70}\text{Mg}_{10}\text{Cu}_{20}$; $T_m = 386^\circ\text{C}$ for $\text{Ca}_{70}\text{Mg}_{20}\text{Cu}_{10}$, $T_m = 363\text{-}364^\circ\text{C}$ for the alloys $\text{Ca}_{65}\text{Mg}_{20}\text{Cu}_{15}$ and $\text{Ca}_{65}\text{Mg}_{25}\text{Cu}_{10}$; $T_m = 374\text{-}377^\circ\text{C}$ for the alloys $\text{Ca}_{40}\text{Mg}_{30}\text{Cu}_{30}$, $\text{Ca}_{40}\text{Mg}_{25}\text{Cu}_{35}$ and $\text{Ca}_{45}\text{Mg}_{19}\text{Cu}_{36}$; and $T_m = \sim 352\text{-}357^\circ\text{C}$ for other 20 alloys. Table 1 also shows that two alloys, $\text{Ca}_{50}\text{Mg}_{25}\text{Cu}_{25}$ and $\text{Ca}_{53}\text{Mg}_{23}\text{Cu}_{24}$, have the lowest melting range, $\Delta T_m = T_l - T_m$, of ~28°C; which indicates the presence of an eutectic reaction at $T_m = \sim 354^\circ\text{C}$, near a composition of ~51 at.% Ca and 24 at.% Mg. Our conclusion that the position of the eutectic shown in Figure 5 is incorrect is also supported by results of Amiya and Inoue [1], who reported $T_m = \sim 370^\circ\text{C}$ for the “eutectic” alloy $\text{Ca}_{67}\text{Mg}_{19}\text{Cu}_{14}$ and $T_m = \sim 357^\circ\text{C}$ for the alloy $\text{Ca}_{57}\text{Mg}_{19}\text{Cu}_{24}$. Detailed assessment of the liquidus projection is in progress in our group and the results will be published shortly [13].

According to the modified eutectic location, the best glass forming alloys, that have the maximum thicknesses in the range of 6 to 10 mm, have near eutectic compositions, which is in agreement with previous experimental observations [11]. The worst glass forming alloys, that have the maximum thickness of ~0.5 mm, are located far from the eutectic and have a very high freezing range (such as $\text{Ca}_{65}\text{Mg}_5\text{Cu}_{30}$ and $\text{Ca}_{55}\text{Mg}_{10}\text{Cu}_{35}$, with $\Delta T_m = \sim 130^\circ\text{C}$ and $\sim 140^\circ\text{C}$, respectively); or high melting temperatures, indicating peritectic reactions (such as $\text{Ca}_{70}\text{Mg}_{20}\text{Cu}_{10}$ and $\text{Ca}_{70}\text{Mg}_{10}\text{Cu}_{20}$, which melting temperatures are about 35°C higher than the eutectic reaction). Both situations should favor crystallization during copper mold casting [5,11]. Analysis of the effect of the alloy composition on the position and width of the X-ray amorphous halo may provide some information on the nearest interatomic spacing in the amorphous

structure [14]. The position of an X-ray halo maximum is known to be directly related with the average radius of the first coordination shell, R^* , through the Bragg equation:

$$R^* = \lambda/(2\sin\Theta^*) \quad (2)$$

where λ is the X-ray wave length (= 0.15418 nm for the Cu K α radiation) and $2\Theta^*$ is the scatter angle at the halo maximum. Figure 6 shows the intensity of the X-ray scattering versus $R = \lambda/(2\sin\Theta)$ for several amorphous alloys $\text{Ca}_{50+x}\text{Mg}_{20}\text{Cu}_{30-x}$, where 2Θ is the angle of deviation of the diffracted beam. It can be seen that the value of R^* corresponding to the position of the scatter intensity peak decreases from ~ 0.3028 nm to ~ 0.2771 nm when the amount of Cu increases from 10 at.% to 30 at.%. This decrease in the average nearest inter-atomic spacing can be explained by a smaller radius of the Cu atoms relative to the radii of Ca and Mg atoms, which should increase the fraction of Ca-Cu and Mg-Cu bonds. The calculated values for the fist coordination shell radius can be compared with the atomic and covalent radii of the elements (Table 2) and atomic and covalent bond distances between Ca-Ca, Ca-Mg, Ca-Cu, and Mg-Cu (Table 3). These comparisons show that the calculated radius values for the first coordination shell can only be obtained if covalent bonds between the elements dominate. An increasing fraction of covalent bonding is known to increase the glass forming ability [15], and this can be one of the reasons of very good glass forming ability of the Ca-based metallic glasses.

CONCLUSIONS

Twenty seven ternary Ca-Mg-Cu bulk metallic glasses were produced by a copper mold casting method and their thermal properties were determined. The critical thickness, τ_{cr} , at which the alloys are fully amorphous after casting is sensitive to the alloy composition and it has a maximum value of at least 10 mm for the alloy $\text{Ca}_{50}\text{Mg}_{22.5}\text{Cu}_{27.5}$. The alloy $\text{Ca}_{50}\text{Mg}_{25}\text{Cu}_{25}$ has $\tau_{cr} = 9$ mm; the alloys $\text{Ca}_{50}\text{Mg}_{20}\text{Cu}_{30}$ and $\text{Ca}_{55}\text{Mg}_{25}\text{Cu}_{20}$ have $\tau_{cr} = 8$ mm; $\text{Ca}_{53}\text{Mg}_{23}\text{Cu}_{24}$ has $\tau_{cr} = 7$ mm; $\text{Ca}_{45}\text{Mg}_{25}\text{Cu}_{30}$, $\text{Ca}_{47.5}\text{Mg}_{22.5}\text{Cu}_{30}$, and $\text{Ca}_{58}\text{Mg}_{18}\text{Cu}_{24}$, have $\tau_{cr} = 6$ mm; $\text{Ca}_{40}\text{Mg}_{25}\text{Cu}_{35}$, $\text{Ca}_{60}\text{Mg}_{20}\text{Cu}_{20}$, and $\text{Ca}_{65}\text{Mg}_{15}\text{Cu}_{20}$ have $\tau_{cr} = 4$ mm; and other alloys have the critical thickness between 0.5 and 3.0 mm. The glass transition temperature for these alloys varies from 83°C to 133°C and the crystallization temperature varies from 112°C to 169°C. Though the glass transition and crystallization temperatures have the maximum values for the best glass forming alloys, no direct dependence of these temperatures on the glass forming ability (or critical thickness) was found for other alloys. Ca-Mg-Cu alloys with good glass forming ability have the

lowest melting temperature ($T_m = 354\text{-}357^\circ\text{C}$) and a rather narrow melting range, indicating that these alloys have near-eutectic compositions. Estimations of the average inter-atomic spacing in the Ca-Mg-Cu metallic glasses from the position of the amorphous halo during X-ray scattering lead to a conclusion that covalent bonds dominate in these glasses.

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TABLES

"Composition Range and Glass Forming Ability of Ternary Ca-Mg-Cu Bulk Metallic Glasses,"
by Senkov and Scott

Table 1. Critical thickness, τ_{cr} , glass transition, T_g , crystallization, T_x , melting, T_m , and liquidus, T_l , temperatures and heats of crystallization, ΔH_x , and fusion, ΔH_m , for Ca-Mg-Zn metallic glasses.

Alloy	τ_{cr} (mm)	T_g (°C)	T_x (°C)	T_m (°C)	T_l (°C)	ΔH_x (J/g)	ΔH_m (J/g)
Ca ₄₀ Mg ₃₀ Cu ₃₀	0.5	122	157	374	421	91	221
Ca ₄₀ Mg ₂₅ Cu ₃₅	4.0	126	163	377	407	89	196
Ca ₄₅ Mg ₃₀ Cu ₂₅	1.0	128	163	354	444	94	241
Ca ₄₅ Mg ₂₅ Cu ₃₀	6.0	127	165	354	405	132	222
Ca ₄₅ Mg ₁₉ Cu ₃₆	0.5	126	155	376	441	106	190
Ca _{47.5} Mg _{22.5} Cu ₃₀	6	126	167	352	400	154	219
Ca ₅₀ Mg ₃₀ Cu ₂₀	2.0	129	166	354	458	104	235
Ca ₅₀ Mg ₂₅ Cu ₂₅	9.0	127	166	354	382	129	210
Ca ₅₀ Mg _{22.5} Cu _{27.5}	10.0	127	169	354	390	150	208
Ca ₅₀ Mg ₂₀ Cu ₃₀	8.0	128	169	355	417	139	378
Ca ₅₃ Mg ₂₃ Cu ₂₄	7.0	133	166	354	382	134	196
Ca ₅₅ Mg ₂₅ Cu ₂₀	8.0	125	155	354	395	125	226
Ca ₅₅ Mg ₂₀ Cu ₂₅	2.0	126	153	355	447	106	170
Ca ₅₅ Mg ₁₅ Cu ₃₀	3.0	124	164	353	433	137	220
Ca ₅₅ Mg ₁₀ Cu ₃₅	0.5	124	149	356	497	98	198
Ca ₅₈ Mg ₁₈ Cu ₂₄	6.0	115	153	355	394	119	186
Ca ₆₀ Mg ₂₅ Cu ₁₅	2.0	117	143	355	403	102	198
Ca ₆₀ Mg ₂₀ Cu ₂₀	4.0	114	139	356	405	105	181
Ca ₆₀ Mg ₁₅ Cu ₂₅	1.0	123	155	354	414	99	205
Ca ₆₀ Mg ₁₃ Cu ₂₇	1.0	121	153	355	428	98	204
Ca ₆₅ Mg ₂₅ Cu ₁₀	0.5	132	156	364	418	96	204
Ca ₆₅ Mg ₂₀ Cu ₁₅	2.0	113	132	363	406	101	193
Ca ₆₅ Mg ₁₅ Cu ₂₀	4.0	110	136	357	409	115	195
Ca ₆₅ Mg ₁₀ Cu ₂₅	2.0	115	147	357	438	114	177
Ca ₆₅ Mg ₅ Cu ₃₀	0.5	130	151	357	484	89	155
Ca ₇₀ Mg ₂₀ Cu ₁₀	0.5	83	112	386	429	87	153
Ca ₇₀ Mg ₁₀ Cu ₂₀	1.0	112	134	397	440	107	175

Table 2. Atomic (r_A) and covalent (r_C) radii of Ca, Mg and Cu.

	r_A nm	r_C nm
Ca	0.197	0.174
Mg	0.160	0.136
Cu	0.128	0.117

Table 3. Atomic (R_A) and covalent (R_C) bond distances between the elements.

	R_A nm	R_C nm
Ca-Ca	0.394	0.348
Ca-Mg	0.357	0.310
Ca-Cu	0.325	0.291
Mg-Mg	0.320	0.272
Mg-Cu	0.288	0.253

FIGURE CAPTIONS

Figure 1. (a) DSC curves of a copper mold cast alloy Ca₅₃Mg₂₃Cu₂₄ of different thicknesses. Locations of glass transition, T_g , and crystallization, T_x , temperatures, as well as the temperatures of the start, T_m , and completion, T_l , of melting, are indicated on one of the curves. An exothermic crystallization reaction with the heat of crystallization ΔH_x is also indicated as an area under the exothermic peaks and above the dashed baseline. Exothermic reactions are up and the heating rate is 40 K/min. (b) Dependences of the heat of crystallization and the volume percent of the amorphous phase on the sample thickness. The critical thickness ($\tau_{cr} = 7$ mm) above which the alloy is partially amorphous after casting is shown.

Figure 2. X-ray diffraction patterns of samples extracted from regions of different thicknesses (shown in the figure) of a copper mold wedge cast alloy Ca₅₃Mg₂₃Cu₂₄.

Figure 3. (a) DSC and (b) X-ray patterns of samples extracted from the regions of different thicknesses (shown in the figures) of the copper mold wedge cast alloys Ca₅₈Mg₁₈Cu₂₄ and Ca₅₀Mg₂₅Cu₂₅.

Figure 4. (a) DSC and (b) X-ray patterns of samples extracted from the regions of critical thicknesses [shown in the figure (a)] of the copper mold wedge cast alloys Ca_{50+x}Mg₂₀Cu_{30-x}, where X = 0, 5, 10, 15, and 20.

Figure 5. Liquidus projection of the Ca-Mg-Cu system, according to [12]. The location of the ternary eutectic (L \rightleftharpoons Ca + CaMg₂ + Ca₂Cu, 67 at. % Ca and 18 at.% Mg) is marked by open triangle and of the three peritectic reactions [(i) L + CaCu₅ \rightleftharpoons CaCu + Cu₂Mg; (ii) L + Cu₂Mg \rightleftharpoons CaCu + CaMg₂; (iii) L + CaCu \rightleftharpoons CaMg₂ + Ca₂Cu] are marked by an open circle. Compositions of 27 amorphous alloys studied in the current work are shown as solid circles. A trapezoid, with a dashed line border, represents a composition range for Ca-Mg-Cu bulk metallic glasses, according to equation (1).

Figure 6. X-ray diffraction intensity vs. inter-atomic spacing in amorphous Ca_{70-x}Mg₂₀Cu_{10+x} alloys (x = 0, 5, 10, 15, and 20).

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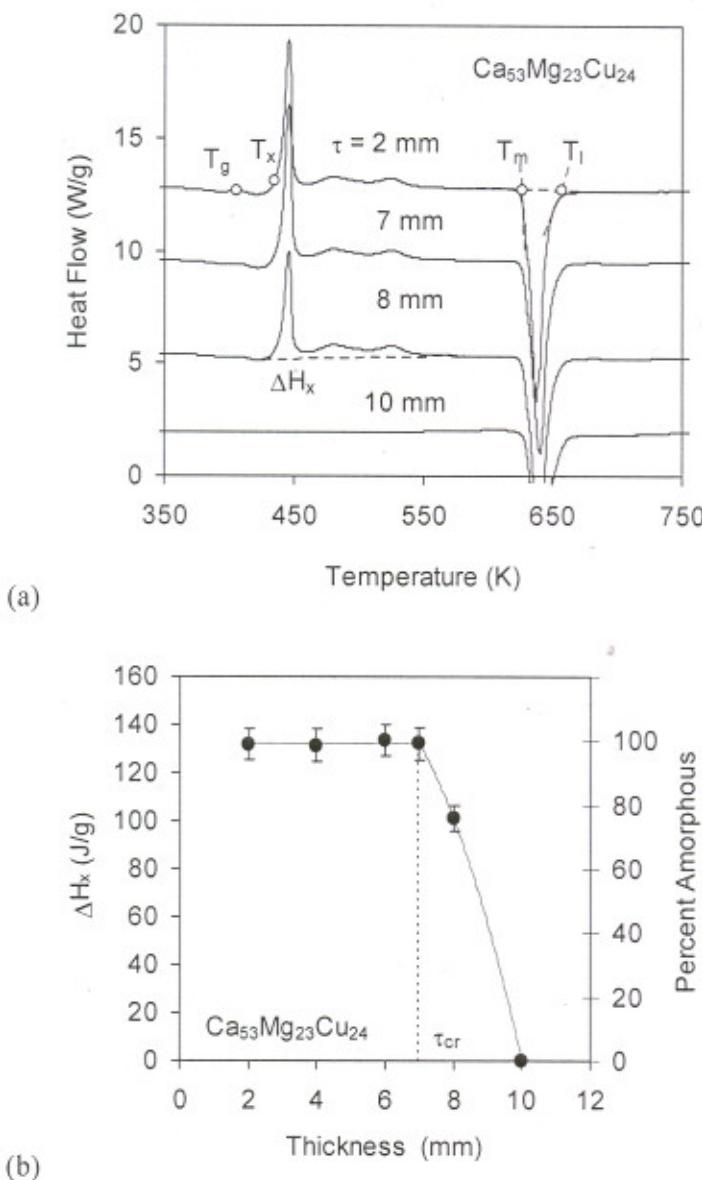


Figure 1. (a) DSC curves of a copper mold cast alloy Ca₅₃Mg₂₃Cu₂₄ of different thicknesses. Locations of glass transition, T_g , and crystallization, T_x , temperatures, as well as the temperatures of the start, T_m , and completion, T_l , of melting, are indicated on one of the curves. An exothermic crystallization reaction with the heat of crystallization ΔH_x is also indicated as an area under the exothermic peaks and above the dashed baseline. Exothermic reactions are up and the heating rate is 40 K/min. (b) Dependences of the heat of crystallization and the volume percent of the amorphous phase on the sample thickness. The critical thickness ($\tau_{cr} = 7$ mm) above which the alloy is partially amorphous after casting is shown.

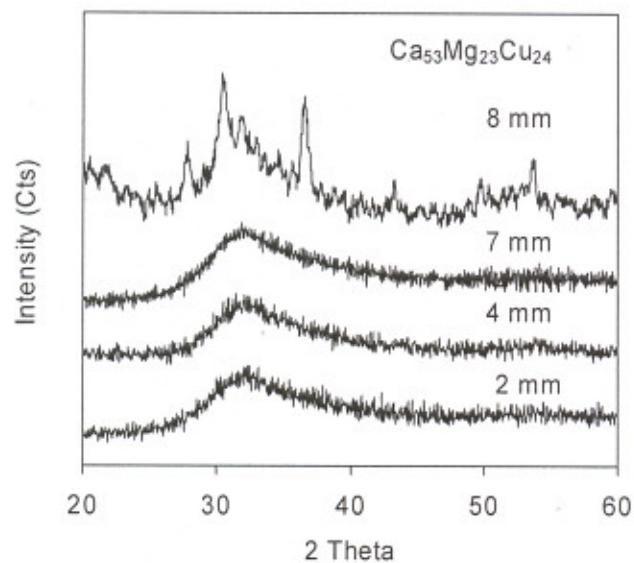


Figure 2. X-ray diffraction patterns of samples extracted from regions of different thicknesses (shown in the figure) of a copper mold wedge cast alloy $\text{Ca}_{53}\text{Mg}_{23}\text{Cu}_{24}$.

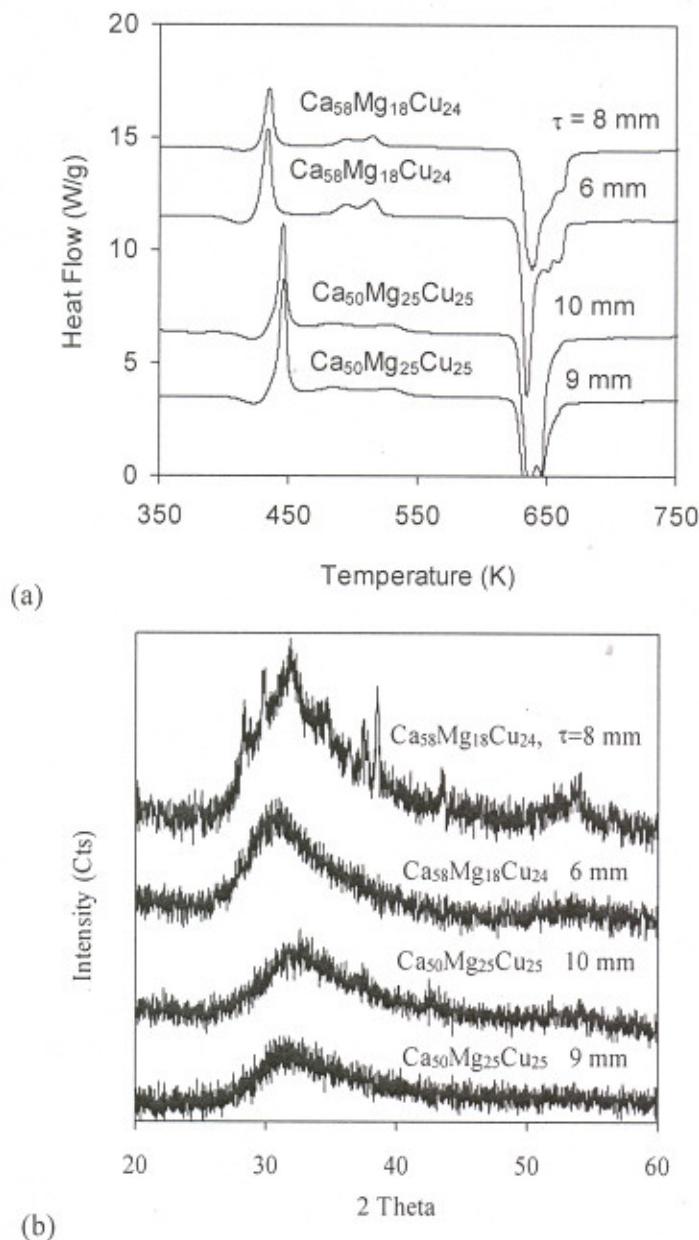
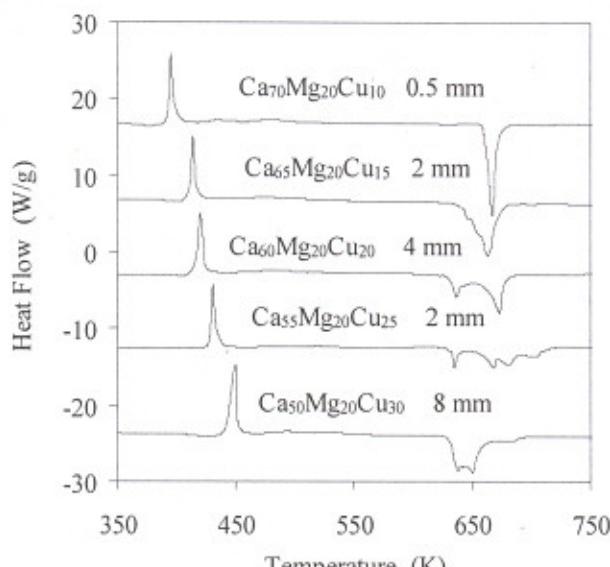
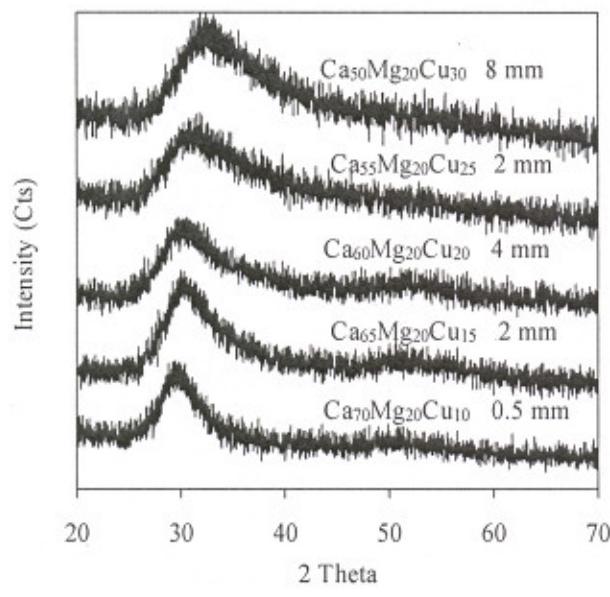


Figure 3. (a) DSC and (b) X-ray patterns of samples extracted from the regions of different thicknesses (shown in the figures) of the copper mold wedge cast alloys $\text{Ca}_{58}\text{Mg}_{18}\text{Cu}_{24}$ and $\text{Ca}_{50}\text{Mg}_{25}\text{Cu}_{25}$.



(a)



(b)

Figure 4. (a) DSC and (b) X-ray patterns of samples extracted from the regions of critical thicknesses [shown in the figures] of the copper mold wedge cast alloys $\text{Ca}_{50+\text{X}}\text{Mg}_{20}\text{Cu}_{30-\text{X}}$, where $\text{X} = 0, 5, 10, 15$, and 20 .

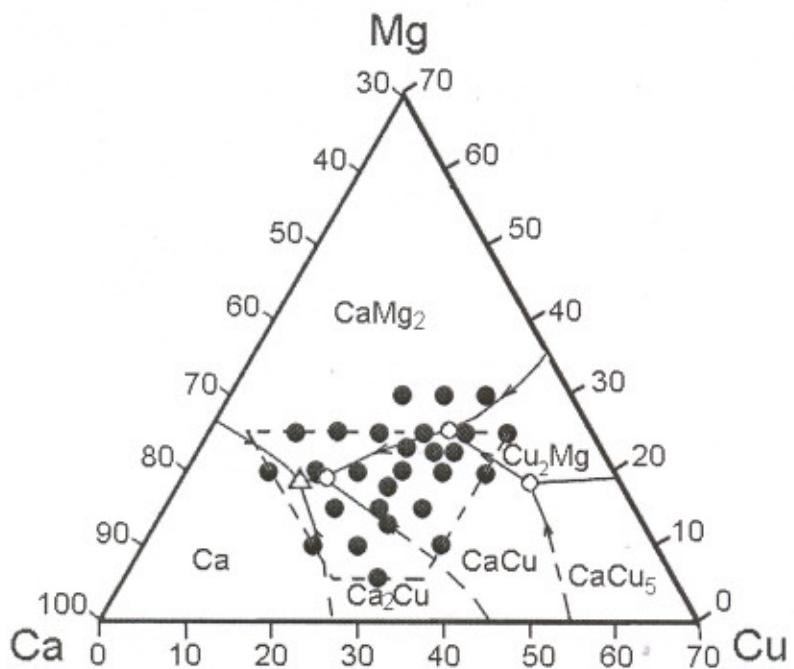


Figure 5. Liquidus projection of the Ca-Mg-Cu system, according to [12]. The location of the ternary eutectic ($L \rightleftharpoons Ca + CaMg_2 + Ca_2Cu$, 67 at. % Ca and 18 at.% Mg) is marked by open triangle and of the three peritectic reactions [(i) $L + CaCu_5 \rightleftharpoons CaCu + Cu_2Mg$; (ii) $L + Cu_2Mg \rightleftharpoons CaCu + CaMg_2$; (iii) $L + CaCu \rightleftharpoons CaMg_2 + Ca_2Cu$] are marked by open circles. Compositions of 27 amorphous alloys studied in the current work are shown as solid circles. A trapezoid, with a dashed line border, represents a composition range for Ca-Mg-Cu bulk metallic glasses, according to equation (1).

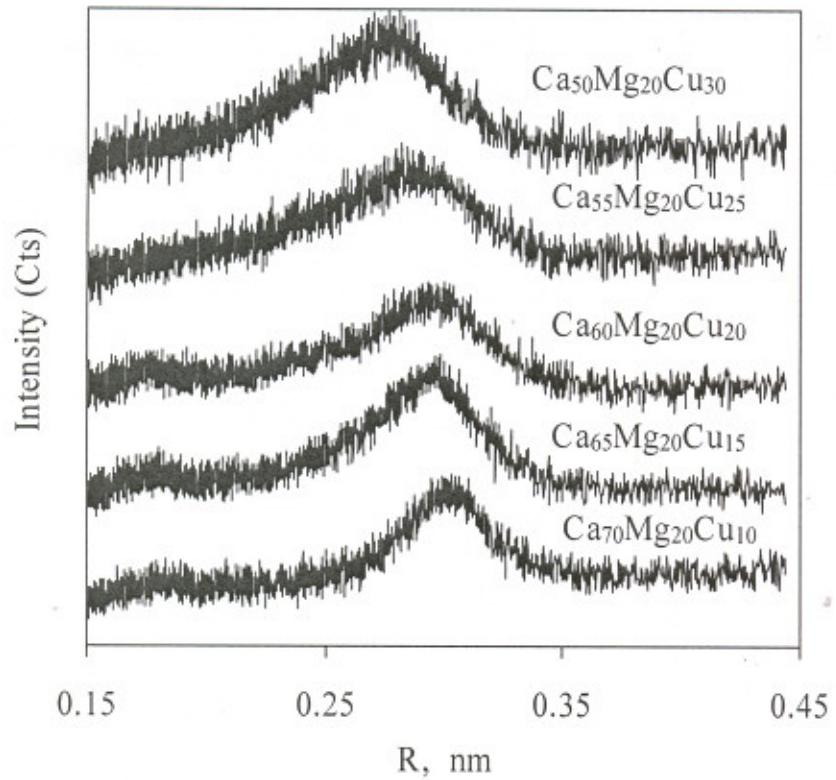


Figure 6. X-ray diffraction intensity vs. inter-atomic spacing in amorphous Ca_{70-x}Mg₂₀Cu_{10+x} alloys ($x = 0, 5, 10, 15$, and 20).